Acetyl-m-oxyphenylacetketodihydrobenzmetoxazine.—This was prepared as in the other cases by treating the above benzylideneanthranilic acid on the water bath with an excess of acetic anhydride. In this case, as in the case of the corresponding *o*- and *p*- compounds described by Ekeley and Dean,[1] the hydrogen of the phenol group has been replaced by acetyl. The crude compound was recrystallized from benzene, yielding colorless crystals. M. p. 124°.

Calc. for $C_{18}H_{15}O_5N$: N, 4.31; found: N, 3.99.

Repeated attempts to obtain oxazines using anisylaldehyde, piperonal, and furfuryl aldehyde have failed. It is not clear wherein the difficulty lies.

BOULDER, COLORADO.

NOTE.

A Case of Potassium Cyanide Poisoning.—Certain data recently obtained in a toxicological investigation seem worth putting on record. A man died from drinking whisky containing an unknown amount of potassium cyanide. The autopsy was performed about six hours after death, and the stomach and intestine placed in glass jars, where they remained unopened until four days after death, when the analysis was commenced.

Potassium cyanide was determined by the well-known method of acidifying with tartaric acid and distilling with steam, adding borax to the distillate and distilling again, and then treating this distillate with silver nitrate to precipitate silver cyanide, which was filtered out, dried, and weighed.

The stomach measured nine and one-half inches along the lesser curvature, and twenty-five and one-half inches along the greater. There were dark areas along the greater curvature, and the wall in places was abnormally red. The rugae were well obliterated. The contents were acid to litmus, and consisted of an undigested meal. The weight of the stomach wall was 288 grams, and of the stomach contents 1014 grams.

A portion of the stomach contents weighing 349 grams, in which hydrocyanic acid was determined five days after death, showed that for the whole stomach contents 0.053 gram potassium cyanide would have been necessary to produce the amount of hydrocyanic acid present. A portion of the intestine analyzed at the same time gave no cyanide. To test the rate of decay of the hydrocyanic acid in the stomach contents, the unused material was placed in a glass stoppered bottle and allowed to stand for several weeks. A portion weighing 143 grams, analyzed 23 days after death, gave 0.039 gram potassium cyanide for the whole stomach contents. A third portion, weighing 310 grams, analyzed 76 days after

[1] THIS JOURNAL, 34, 164.

death, showed no hydrocyanic acid present. The material showed no marked decomposition.

A portion of the stomach wall, about one square inch in area, taken when the analysis was begun four days after death, extracted with 0.4% hydrochloric acid gave a solution which did not dissolve fibrin. A solution from a stomach wall where death has been due to natural causes will ordinarily do so after that length of time. M. P. CRAM.

BOWDOIN COLLEGE, BRUNSWICK, MAINE.

CORRECTION.

THE MOLECULAR REARRANGEMENT OF TRIARYLMETHYLHYDROXYLAMINES AND THE "BECKMANN" REARRANGEMENT OF KETOXIMES.

BY JULIUS STIEGLITZ AND PAUL NICHOLAS LEECH.

Through an oversight in the final editing we have given (Vol. XXXVI, p. 297) the proportions by weight of aniline hydrochloride and chloroaniline hydrochloride obtained in the rearrangement of parachlorophenyldiphenylmethylhydroxylamine. The molal ratios are:

Line 24, 28% chloroaniline, 72% aniline.

Line 29, 31.5% chloroaniline, 68.5% aniline.

Line 30, 28% chloroaniline, 72% aniline.

The provisional conclusion based on these results (p. 287), that "roughly two-thirds" of the substance rearranges to form aniline and one-third to form chloroaniline is little affected by this correction.

CHICAGO, Feb. 7, 1914.

NEW BOOKS.

A Textbook of Chemistry. By WILLIAM A. NOYES, Director of the Chemical Laboratory of the University of Illinois. Henry Holt & Co., New York, 1913. xv + 602 pp. 13.5 × 20.5 cm. Price, $2.25.

This new text by Professor Noyes marks a decided step forward in the presentation of the facts and theories of general chemistry. The brief introduction and very early entrance upon equations (p. 21) are special features of the book. The usual discussions on the gas laws follow in close order with good illustrations, but, in the minds of the great majority of chemists, it is doubtful if Avogadro's hypothesis (here called a law) would be construed as other than an hypothesis. It is pleasing to note the absence of the term "equivalent weight," a term that has too long outlived its usefulness and only served for the confusion of the beginner. Graphical formulas meet with an early introduction, simplifying remarkably such discussions as the Weldon process for chlorine. The chapter on the Periodic System is introduced just after chlorine and preceding the chapter on the halogen family (Chapter X). It is surprising